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Assessment page 121 20. Explain how the type of an atom is defined. by the atomic number 21. Chapter 4 Solutions Manual Chemistry

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observation that mass is conserved in
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because atoms cannot be created,
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100 atoms of gallium there are: 60
atoms of mass 69 = $60 \times 69 = 4140$
40 atoms of mass 71 = $40 \times 71 = 2840$
----- Total
mass of 100 atoms = 6980
Average mass of 1 atom = 69.8.

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4.1. For the reaction $R \rightarrow P$, the concentration of a reactant changes from 0.03 M to 0.02 M in 25 minutes. Calculate the average rate of reaction using units of time both in minutes and seconds.

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