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4.2 For the reaction, 2A + B → A<sub>2</sub>B, the rate = k<sub>1</sub>[A]<sup>2</sup>[B] with k = 2.0 × 10<sup>-6</sup> mol<sup>-2</sup> L<sup>2</sup> s<sup>-1</sup>. Calculate the initial rate of the reaction when [A] = 0.1 mol L<sup>-1</sup>, [B] = 0.2 mol L<sup>-1</sup>. Calculate the rate of reaction after [A] is reduced to 0.06 mol L<sup>-1</sup>. Sol. Initial rate of reaction = k [A]<sup>2</sup> [B] 2 = (20 × 10<sup>-6</sup> mol<sup>-2</sup> s<sup>-1</sup>) (0.1 mol L<sup>-1</sup>)<sup>2</sup> (0.2 mol L<sup>-1</sup>) 2 = 8 × 10<sup>-9</sup> mol<sup>-1</sup> s<sup>-1</sup>.

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Solutions Manual Chemistry: Matter and Change • Chapter 11 211 CHAPTER 11 SOLUTIONS MANUAL Section 11.1 Assessment page 372 5. Compare the mass of the reactants and the mass of the products in a chemical reaction, and explain how these masses are related. The coefficients in the balanced equation indicate the molar relationship between each ...