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Chemistry Chapter 14 Notes. reaction rate. reaction rate equation. instantaneous rate. General equation. the rate at which reactants change into products over time. change in concentration/change in time. The reaction rate at a particular time as opposed to the average rate.  $aA + bB = cC + dD$ .

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14.1 465 McNary-Haugen Bill Bill that would have provided federal price-supports which would have helped stabilized prices in domestic market, however, the bill was vetoed by President Coolidge and failed to ever pass. 14.1 465 price-supports

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This is a case in which the solution contains a mixture of acids of different ionization strengths. In solution, the  $\text{HCO}_2\text{H}$  exists primarily as  $\text{HCO}_2\text{H}$  molecules because the ionization of the weak acid is suppressed by the strong acid. Therefore, the  $\text{HCO}_2\text{H}$  contributes a negligible amount of hydronium ions to the solution. The stronger acid,  $\text{HCl}$ , is the dominant producer of hydronium ions ...

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